## NONSTOICHIOMETRIC OXIDE FORMATION IN SYSTEM Nd-Ni-O.

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The non-stoichiometric compositions  $Nd_{2-x}NiO_{4-3x/2+\delta}$  were obtained using coprecipitation method from nitrates solutions by  $K_2CO_3$  followed by calcination of the obtained blend at 900°C (10 hours). It was found that full coprecipitation of ions begins at mole ratio of coprecipitated ions to precipitator n=1,75 in range pH=10-10,2.

The thermogravimetric investigations established that coprecipitated carbonates decompose in the four stages.

The first –  $(25 - 392^{\circ}C)$ , removing of crystallisation and adsorbed water evaporation. The second -  $(392 - 520^{\circ}C)$ , OH<sup>-</sup> - groupes decomposition. The third (622 - 720 °C), carbonates groupes decomposition. At 850°C - neodymium nickelate phases formation.

According to X-ray phase analysis  $Nd_2NiO_{4.392(5)}$  consists of orthorombic neodymium nickelate and a small of  $Nd_2O_3$  impurity (specific reflex at d=0,3026 nm).

 $Nd_2NiO_{4.392(5)}$  unit cell parameters - a=0,545(0)nm, b=0,536(8) nm, c=0,123(3)nm, space group Bmab.

Samples  $Nd_{1.9}NiO_{4.113(5)}$  and  $Nd_{1.8}NiO_{4.006(5)}$  consist of non-stoichiometric phase  $Nd_{2-x}NiO_{4+\delta}$  with neodymium ions deficient in sublattice.

The orthorhombic unit cell parameters for  $Nd_{1.9}NiO_{4.113(5)}$  - a=0,541(9) nm, b=0,535(4) nm, c=0,123(2) nm, Nd\_{1.8}NiO\_{4.006(5)} - a=0,541(5) nm, b=0,535(4) nm, c=0,123(2) nm.

The stabilization of non-stoichiometric neodymium nickelate phases likely is due to ion vacancies formation.

 $Nd_{2-x}NiO_{4-3x/2+\delta}$  samples with 0,2<x≤0,5 are polyphase and contain  $Nd_2NiO_4$ , NiO and  $Nd_2O_3$ . It was not observed  $Nd_3Ni_2O_7$  phase formation in our experimental conditions.